Status: Currently Official on 18-Feb-2025
Official Date: Official as of 01-Jan-2018
Document Type: USP Monographs
DocId: GUID-0A0D26B9-A8E8-4938-89AA-7BA6F90AD986_3_en-US
DOI: https://doi.org/10.31003/USPNF_M1308_03_01
DOI Ref: 3s5ro

Sodium Bromide

NaBr 102.89

Sodium bromide.

Sodium bromide CAS RN®: 7647-15-6 ; UNII: LC1V549NOM.

» Sodium Bromide contains not less than 98.0 percent and not more than 100.5 percent of NaBr, calculated on the dried basis. It contains no added substances.

Packaging and storage—Preserve in well-closed containers, and store at room temperature.

Appearance of solution: clear and colorless.

Test solution—Dissolve 10.0 g in carbon dioxide-free water, and dilute with the same solvent to 100 mL.

Identification—

A: A solution containing 4.0 mg of sodium bromide responds to the test for *Bromide* ⟨191⟩.

B: It responds to the test for *Sodium* ⟨191⟩.

Acidity or alkalinity—To 10 mL of the solution prepared for the test for *Appearance of solution,* add 0.1 mL of bromothymol blue TS: not more than 0.5 mL of 0.01 N hydrochloric acid or 0.01 N sodium hydroxide is required to change the color of this solution.

Loss on drying ⟨731⟩—Dry it at 100° to 105° for 3 hours: it loses not more than 3.0% of its weight.

Bromates—

Starch–mercuric iodide solution—Triturate 1.0 g of soluble starch with 5 mL of water, and pour the mixture into 100 mL of boiling water containing 10 mg of mercuric iodide.

Procedure—To 10 mL of the solution prepared for the test for *Appearance of solution* add 1 mL of *Starch–mercuric iodide solution,* 0.1 mL of a 100 g per L solution of potassium iodide, and 0.25 mL of 0.5 M sulfuric acid. Allow to stand protected from light for 5 minutes. No blue or violet color develops.

Limit of chlorine: not more than 0.6%.

Nitric acid solution and *Ferric ammonium sulfate solution*—Proceed as directed in the *Assay.*

Procedure—Dissolve 1.000 g Sodium Bromide in 20 mL of *Nitric acid solution* in a conical flask, add and mix 5 mL of 30 percent hydrogen peroxide, and heat in a water bath until the solution is colorless. Rinse the sides of the flask with a small quantity of water, and heat in a water bath for 15 minutes. Allow to cool, dilute with water to 50 mL, and add 5.0 mL of silver nitrate VS and 1 mL of dibutyl phthalate. Mix, and back titrate the excess silver nitrate with ammonium thiocyanate VS (see *Titrimetry* ⟨541⟩), using 5 mL of *Ferric ammonium sulfate solution* as the indicator. Perform a blank titration. Not more than 1.7 mL of silver nitrate VS is used.

Iodides—To 5 mL of the solution prepared for the test for *Appearance of solution* add 0.15 mL of a 10.5 g per 100 mL ferric chloride solution, and 2 mL of dichloromethane. Shake, and allow to separate. The lower layer is colorless.

Sulfates ⟨221⟩—A 2.0-g portion shows no more sulfate than corresponds to 0.2 mL of 0.020 N sulfuric acid (0.01%).

Limit of iron: not more than 20 ppm.

Citric acid solution—Prepare a 200 mg citric acid per mL solution.

Iron standard solution—Transfer 0.863 g of ferric ammonium sulfate to a 500-mL volumetric flask, and dissolve in 25 mL of dilute sulfuric acid. Dilute with water to volume. Transfer 1.0 mL of the resulting solution to a 10-mL volumetric flask, and dilute with water to volume. Transfer 2.5 mL of this resulting solution to a 50-mL volumetric flask, and dilute with water to volume.

[Note—Prepare immediately before use.]

Test solution—Transfer 5 mL of the solution prepared for the test for *Appearance of solution* to a 10-mL volumetric flask, and dilute with water to volume.

Procedure—To 10 mL each of the *Iron standard solution* and the *Test solution* add 2.0 mL of the *Citric acid solution* and 0.1 mL of thioglycolic acid. Make alkaline to litmus with ammonia water, and dilute with water to 20 mL. After 5 minutes, any pink color in the *Test solution* is not more intense than that in the *Iron standard solution*.

Magnesium and alkaline-earth metals—To 200 mL of water add 0.1 g of hydroxylamine hydrochloride, 10 mL of pH 10.0 ammonia–ammonium chloride buffer (prepared by dissolving 5.4 g of ammonium chloride in 20 mL of water, adding 20 mL of ammonium hydroxide, and diluting to 100 mL), 1 mL of 0.1 M zinc sulfate, and about 0.2 g of eriochrome black T trituration. Heat to about 40°. Titrate this solution (see *Titrimetry* ⟨541⟩) with 0.01 M edetate disodium VS until the violet color changes to deep blue. To this solution add 10.0 g of Sodium Bromide dissolved in 100 mL of water. If the color changes to violet, titrate the solution with 0.01 M edetate disodium VS to a deep blue endpoint. The volume of 0.01 M edetate disodium consumed in the second titration does not exceed 5.0 mL (0.02%, calculated as Ca).

Assay—

Nitric acid solution—Dilute 14 mL of nitric acid with water to 100 mL.

Ferric ammonium sulfate solution—Transfer 10 g of ferric ammonium sulfate to a 100-mL volumetric flask. Dissolve in and dilute with water to volume.

Procedure—Dissolve 2.000 g of Sodium Bromide in water, and dilute with water to 100.0 mL. To 10.0 mL of the solution add 50 mL of water, 5 mL of *Nitric acid solution,* 25.0 mL of silver nitrate VS, and 2 mL of dibutyl phthalate. Mix, and back titrate the excess silver nitrate with ammonium thiocyanate VS (see *Titrimetry* ⟨541⟩), using 2 mL of *Ferric ammonium sulfate solution* as the indicator, shaking vigorously towards the endpoint. Each mL of 0.1 M silver nitrate is equivalent to 10.29 mg of NaBr. Calculate the percent content of Sodium Bromide, corrected for the chloride content, by the formula:

$$a - 2.902b$$

in which a is the percent content of NaBr and NaCl obtained, calculated as NaBr; and b is the percent content of chlorides.

Auxiliary Information - Please check for your question in the FAQs before contacting USP.

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Chromatographic Database Information: Chromatographic Database

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Pharmacopeial Forum: Volume No. PF 31(2)

Current DocID: GUID-0A0D26B9-A8E8-4938-89AA-7BA6F90AD986_3_en-US
Previous DocID: GUID-0A0D26B9-A8E8-4938-89AA-7BA6F90AD986_1_en-US
DOI: https://doi.org/10.31003/USPNF_M1308_03_01
DOI ref: 3s5ro