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Change to read:

0.1 N Potassium Dichromate VS

▲ (USP 1-May-2021)

Dissolve about 5 g of [potassium dichromate](#) in 1000 mL of [water](#).

Standardization

▲ See [Volumetric Solutions, 1. Introduction](#).

See [Titrimetry \(541\)](#).

Standardize by one of the following procedures. [NOTE—Other standardization procedures may be used. See [Volumetric Solutions, 2. Preparation and Standardization, 2.3 Standardization](#).]

Standardization with visual endpoint: ▲ (USP 1-May-2021) Transfer 25.0 mL of this solution to a glass-stoppered, 500-mL flask, add 2 g of [potassium iodide](#). ▲ (USP 1-May-2021) dilute with 200 mL of [water](#), add 5 mL of [hydrochloric acid](#), allow to stand for 10 min in a dark place, and titrate the liberated iodine with [0.1 N sodium thiosulfate VS](#), adding 3 mL of [starch TS](#) as the endpoint is approached. Carry out a blank determination.

$$N = \frac{[\text{mL (sample)} - \text{mL (blank)}] \text{ Na}_2\text{S}_2\text{O}_3 \times N \text{ Na}_2\text{S}_2\text{O}_3}{25.0}$$

Standardization with potentiometric endpoint: Transfer 10.0 mL of this solution to a glass-stoppered, 500-mL flask, add 2 g of [potassium iodide](#), dilute with 200 mL of [water](#), add 5 mL of [hydrochloric acid](#), allow to stand for 10 min in a dark place, and titrate the liberated iodine with [0.1 N sodium thiosulfate VS](#) potentiometrically using a combined platinum (Pt) electrode. Carry out a blank determination.

$$N = \frac{[\text{mL (sample)} - \text{mL (blank)}] \text{ Na}_2\text{S}_2\text{O}_3 \times N \text{ Na}_2\text{S}_2\text{O}_3}{10.0}$$

▲ (USP 1-May-2021)

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.1 N POTASSIUM DICHROMATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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