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0.1 N Potassium Bromate VS

KBrO_3 , 167.00

2.784 g in 1000 mL

Dissolve 2.784 g of [potassium bromate](#) in [water](#) to make 1000 mL.

Standardization: Transfer an accurately measured volume of about 40 mL of the solution to a glass-stoppered flask, add 3 g of [potassium iodide](#), and follow with 3 mL of [hydrochloric acid](#). Allow to stand for 5 min, then titrate the liberated iodine with [0.1 N sodium thiosulfate VS](#), adding 3 mL of [starch TS](#) as the endpoint is approached. Correct for a blank run on the same quantities of the same reagents, and calculate the normality.

$$N = \frac{\text{mL Na}_2\text{S}_2\text{O}_3 \times N \text{ Na}_2\text{S}_2\text{O}_3}{\text{mL KBrO}_3 \text{ Solution}}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.1 N POTASSIUM BROMATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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