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0.05 M Potassium Iodate VS

KIO_3 , 214.00
 10.70 g in 1000 mL

Dissolve 10.700 g of [potassium iodate](#), previously dried at 110° to constant weight, in [water](#) to make 1000.0 mL.

Standardization: To 15.0 mL of solution in a 250-mL iodine flask, add 3 g of [potassium iodide](#) and 3 mL of [hydrochloric acid](#) previously diluted with 10 mL of [water](#). Stopper immediately, and allow to stand in the dark for 5 min. Then add 50 mL of cold [water](#), and titrate the liberated iodine with freshly standardized [0.1 N sodium thiosulfate](#). Add 3 mL of [starch indicator solution](#) near the end of the titration, and continue to the absence of the blue-starch-iodine complex.

$$M = \frac{\text{mL} \times N \text{ Na}_2\text{S}_2\text{O}_3}{\text{mL KIO}_3 \times 6}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.05 M POTASSIUM IODATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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