

Status: Currently Official on 16-Feb-2025  
 Official Date: Official as of 01-May-2017  
 Document Type: Reagents  
 DocId: GUID-996B72F9-6B58-4FF4-9604-AB5F4B9350F9\_2\_en-US  
 DOI: [https://doi.org/10.31003/USPNF\\_R3386\\_02\\_01](https://doi.org/10.31003/USPNF_R3386_02_01)  
 DOI Ref: 16t4s

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## 0.05 M Potassium Ferricyanide VS

$K_3Fe(CN)_6$ , 329.24  
 16.46 g in 1000 mL

Dissolve about 17 g of [potassium ferricyanide](#) in [water](#) to make 1000 mL.

**Standardization:** Transfer 50.0 mL of this solution to a glass-stoppered, 500-mL flask, dilute with 50 mL of [water](#), add 10 mL of [potassium iodide TS](#) and 10 mL of [dilute hydrochloric acid](#), and allow to stand for 1 min. Then add 15 mL of [zinc sulfate](#) solution (1 in 10), and titrate the liberated iodine with [0.1 N sodium thiosulfate VS](#), adding 3 mL of [starch TS](#) as the endpoint is approached.

Protect from light, and restandardize before use.

$$M = \frac{\text{mL Na}_2\text{S}_2\text{O}_3 \times N \text{ Na}_2\text{S}_2\text{O}_3}{50.0}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

**Auxiliary Information** - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.05 M POTASSIUM FERRICYANIDE VS	<a href="#">Margareth R.C. Marques</a> Principal Scientific Liaison	HDQ Headquarters

**Most Recently Appeared In:**

Pharmacopeial Forum: Volume No. PF 42(1)

**Current DocID:** [GUID-996B72F9-6B58-4FF4-9604-AB5F4B9350F9\\_2\\_en-US](#)

**Previous DocID:** [GUID-996B72F9-6B58-4FF4-9604-AB5F4B9350F9\\_1\\_en-US](#)

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**DOI ref:** [16t4s](#)