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Change to read:

0.1 N Cupric Nitrate VS

Dissolve 23.3 g of [cupric nitrate 2.5 hydrate](#), or 24.2 g of the [trihydrate](#), in [water](#) to make 1000 mL.

Standardization

See [▲ *Volumetric Solutions, 1. Introduction*](#).▲ (USP 1-May-2024)

See [Titrimetry \(541\)](#).

Standardize by the following procedure. [NOTE—Other standardization procedures may be used. See [Volumetric Solutions, 2. Preparation and Standardization, 2.3 Standardization](#).]

▲ **Standardization with potentiometric endpoint:**▲ (USP 1-May-2024) Transfer 20.0 mL of the solution to a 250-mL beaker. Add 2 mL of 5 M sodium nitrate, 20 mL of [ammonium acetate TS](#), and sufficient [water](#) to make 100 mL. Titrate with [0.05 M edetate disodium VS](#). Determine the endpoint potentiometrically using a copper ion-selective electrode. Perform a blank determination, and make any necessary correction.▲

$$N = \frac{(\text{mL sample} - \text{mL blank}) \text{ Edetate disodium} \times M \text{ Edetate disodium}}{20.0}$$

▲ (USP 1-May-2024)

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.1 N CUPRIC NITRATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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