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0.1 M Mercuric Nitrate VS

$\text{Hg}(\text{NO}_3)_2$, 324.60
 32.46 g in 1000 mL

Dissolve about 35 g of [mercuric nitrate](#) in a mixture of 5 mL of [nitric acid](#) and 500 mL of [water](#), and dilute with [water](#) to 1000 mL.

Standardization: Transfer an accurately measured volume of about 20 mL of the solution to a conical flask, and add 2 mL of [nitric acid](#) and 2 mL of [ferric ammonium sulfate TS](#). Cool to below 20°, and titrate with [0.1 N ammonium thiocyanate VS](#) to the first appearance of a permanent brownish color.

$$M = \frac{\text{mL NH}_4\text{SCN} \times N \text{ NH}_4\text{SCN}}{\text{mL Hg}(\text{NO}_3)_2 \times 2}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

| Topic/Question | Contact | Expert Committee |
|---------------------------|--|------------------|
| 0.1 M MERCURIC NITRATE VS | Margareth R.C. Marques Principal Scientific Liaison | HDQ Headquarters |

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