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0.1 M Lead Perchlorate VS

$\text{Pb}(\text{ClO}_4)_2 \cdot 3\text{H}_2\text{O}$, 460.15
 46.01 g in 1000 mL

Dissolve 46 g of [lead perchlorate](#) in [water](#), and dilute with [water](#) to 1000.0 mL.

Standardization: Accurately weigh about 150 mg of [sodium sulfate](#), previously dried at 105° for 4 h, and dissolve in 50 mL of [water](#). Add 50 mL of a mixture of [water](#) and [formaldehyde](#) (1:1), and stir for about 1 min. Determine the endpoint potentiometrically using a lead ion selective electrode. Perform a blank determination, and make any necessary corrections. Each 14.204 mg of sodium sulfate is equivalent to 1 mL of 0.1 M lead perchlorate.

$$M = \frac{\text{mg sodium sulfate}}{142.04 \times \text{mL lead perchlorate}}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.1 M LEAD PERCHLORATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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