Add the following:

▲0.01 N Silver Nitrate VS

Transfer 100 mL of 0.1 N silver nitrate VS to a 1000-mL volumetric flask. Dilute with water to volume. Otherwise, dissolve about 1.75 g of silver nitrate in 1000 mL of water.

Standardization

See *Volumetric Solutions, 1. Introduction*.

See *Titrimetry ⟨541⟩*.

Standardize by one of the following procedures. [NOTE—Other standardization procedures may be used. See *Volumetric Solutions, 2. Preparation and Standardization, 2.3 Standardization*.]

Standardization with visual endpoint: Transfer 0.50 mL of 0.1 N sodium chloride VS to a 150-mL beaker containing 50 mL of water, and add 50 mL of alcohol, 5 mL of acetic acid, and 0.5 mL of eosin Y TS. Stir, preferably with a magnetic stirrer, and titrate with the silver nitrate solution.

$$N = \frac{mL\ NaCl \times N\ NaCl}{mL\ AgNO_3}$$

Standardization with potentiometric endpoint: Transfer 0.50 mL of 0.1 N sodium chloride VS to a 150-mL beaker containing 50 mL of water, and add 50 mL of alcohol, 5 mL of acetic acid. Stir, preferably with a magnetic stirrer, and titrate with the silver nitrate solution using a combined silver electrode.

$$N = \frac{mL\ NaCl \times N\ NaCl}{mL\ AgNO_3}$$

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]▲ (USP 1-AUG-2023)

Auxiliary Information - Please check for your question in the FAQs before contacting USP.

Topic/Question	Contact	Expert Committee
0.01 N SILVER NITRATE VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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