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Change to read:

0.01 M Edetate Disodium VS

▲ (USP 1-Aug-2024)

Transfer 100 mL of [0.1 M edetate disodium VS](#) to a 1000-mL volumetric flask. Dilute with [water](#) to volume.

Standardization

▲ See [Volumetric Solutions, 1. Introduction](#).

See [Titrimetry \(541\)](#).

Standardize by one of the following procedures. [NOTE—Other standardization procedures may be used. See [Volumetric Solutions, 2. Preparation and Standardization, 2.3 Standardization](#).]

Standardization with visual endpoint: ▲ (USP 1-Aug-2024) Accurately weigh about 40 mg of [chelometric standard calcium carbonate](#), previously dried at ▲210 ▲ (USP 1-Aug-2024) ° for ▲4 ▲ (USP 1-Aug-2024) h and cooled in a desiccator, or dried according to the label instructions, transfer to a 400-mL beaker, add 10 mL of [water](#), and swirl to form a slurry. Cover the beaker with a watch glass, and introduce 4 mL of [diluted hydrochloric acid](#) from a pipet inserted between the lip of the beaker and the edge of the watch glass. Wash down the sides of the beaker, the outer surface of the pipet, and the watch glass with [water](#), and dilute with [water](#) to about 200 mL. While stirring the solution, preferably with a magnetic stirrer, add about 30 mL of the edetate disodium solution from a 50-mL buret. Adjust the solution with [sodium hydroxide TS](#) to a pH of 12–13, add 300 mg of [hydroxy naphthol blue](#), and continue the titration with edetate disodium solution to a blue endpoint.

$$▲M = \frac{(g \text{ CaCO}_3)(1000)(\text{Assay}/100)}{100.09 \times \text{mL EDTA}}$$

where Assay is the content/potency of calcium carbonate.

Standardization with potentiometric endpoint: Accurately weigh about 100 mg of [chelometric standard calcium carbonate](#), previously dried at 210° for 4 h and cooled in a desiccator, or dried according to the label instructions, transfer to a 100-mL volumetric flask, add 10 mL of [water](#), and swirl to form a slurry. Introduce [diluted hydrochloric acid](#) from a pipet until the calcium carbonate is completely dissolved, then dilute with [water](#) to volume. Transfer 10.0 mL of the above calcium carbonate solution into a 200-mL beaker and dilute with [water](#) to about 100 mL. Add 5 mL of [ammoniacal buffer, pH 10](#), and titrate potentiometrically with the edetate disodium solution using a combined calcium ion-selective electrode.

$$M = \frac{(g \text{ CaCO}_3)(1000)(\text{Assay}/100) \times 0.1}{100.09 \times \text{mL EDTA}}$$

where Assay is the content/potency of calcium carbonate. ▲ (USP 1-Aug-2024)

[NOTE—If this volumetric solution is used in a qualitative application such as pH adjustment, dissolution medium, or diluent, its standardization is not required.]

Auxiliary Information - Please [check for your question in the FAQs](#) before contacting USP.

Topic/Question	Contact	Expert Committee
0.01 M EDETATE DISODIUM VS	Margareth R.C. Marques Principal Scientific Liaison	HDQ Headquarters

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